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Solution: OH^- OH^- OH^- 0.50 mmol mL $n = \frac{m}{M} = \frac{0.50 \text{ mmol}}{0.10 \text{ mL}} = 5.0 \text{ mmol/L}$ HCl(aq) HCl(aq) $[\text{HCl(aq)}]$ 12.5 mmol = $n \times V$ 0.10 mmol HCl(aq) /mL $V = 120 \text{ mL}$ 2. (a) Given: $V \text{ HCl(aq)} = 25.00 \text{ mL}$; $[\text{HCl(aq)}] = 0.350 \text{ mol/L}$; $V \text{ NaOH(aq)} = 5.0 \text{ mL}$; $[\text{NaOH(aq)}] = 0.500 \text{ mol/L}$; Required: $[\text{H}^+(\text{aq})]$, pH
Analysis: Before titration, amount of H^+ is: $\text{H}^+(\text{aq})$ $\text{H}^+(\text{aq})$ $\text{H}^+(\text{aq})$ 0.350 mmol mL $n = \frac{m}{M} = \frac{0.350 \text{ mmol}}{0.10 \text{ mL}} = 3.5 \text{ mmol/L}$

~~Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547~~

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8 MHR Chemistry 12 Solutions Manual 978-0-07-106042-4 8. Identify any errors in the structure by drawing them. Rename the structure correctly. 2-ethyl-2,4,4-trimethylpentane What Is Required? You are to identify the errors in the name by drawing the structure and renaming it correctly. What Is Given? You are given an incorrectly named alkane.

~~Unit 1 Organic Chemistry - Mr. Arthur's Science Page~~

= 0.12 mol/L Statement: The equilibrium concentration of dinitrogen tetroxide gas is 0.12 mol/L. 2. Given: initial quantity of $\text{NOCl(g)} = 3.00 \text{ mol}$; volume = 3.00 L; $[\text{NO(g)}]$ initial = 0.00 mol/L; $[\text{Cl}_2(\text{g})]$ initial = 0.00 mol/L; $[\text{NO(g)}]$ equilibrium = 0.043 mol/L Required: $[\text{NOCl(g)}]$ equilibrium; $[\text{Cl}_2(\text{g})]$ equilibrium Solution: Step 1.

~~Section 7.1: Equilibrium Systems Mini Investigation: The ...~~

Solution: bonds broken 3 mol 1 mol CH_3Cl 3 mol H_2 D_2 D_2 $\Delta H = 413 \text{ kJ mol}^{-1} + 1 \text{ mol} \times 339 \text{ kJ mol}^{-1} = 1578 \text{ kJ}$ Statement: The energy required to separate 1 mol of chloromethane into free atoms is 1578 kJ. 2. Solution: Step 1: Write the balanced chemical equation for the combustion of 1 mol of ethene gas. $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$

~~Section 5.3: Bond Energies Tutorial 1 Practice, page 312~~

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Solution: $q = Ne = (1.2 \times 10^8)(1.602 \times 10^{-19} \text{ C}) = 1.92 \times 10^{11} \text{ C}$ (one extra digit carried) $F = kq_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N} \cdot \text{m}^2 / \text{C}^2 \times (1.92 \times 10^{11} \text{ C})^2 / (1.10 \text{ m})^2 = 2.7 \times 10^{12} \text{ N}$ Statement: The force of repulsion between the two plastic spheres is $2.7 \times 10^{12} \text{ N}$. 2. Given: $m = 2.48 \times 10^{-15} \text{ kg}$; $d = 1.7 \text{ cm} = 0.017 \text{ m}$; $V = 260 \text{ V}$; $e = 1.602 \times 10^{-19} \text{ C}$; $g = 9.8 \text{ m/s}^2$ Required: q ; N

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